

Descriptor-Based Analysis Applied to HCN Synthesis from NH₃ and CH₄**

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The design of solid metal catalysts by theoretical methods has been a long-standing goal in heterogeneous catalysis.^[1,2] Recent developments in methodology and computer technology, as well as the establishment of a descriptor-based approach for the analysis of reaction mechanisms and trends across the periodic table, allow fast screening for new catalytic materials and have led to the first examples of computational discoveries of new materials.^[3–8] The underlying principles of the descriptor-based approach are the existence of relations between the surface electronic structure, adsorption energies, and activation barriers that result in volcano-shaped activity plots as a function of simple descriptors, such as atomic binding energies or the d-band center.^[1,6,9–17]

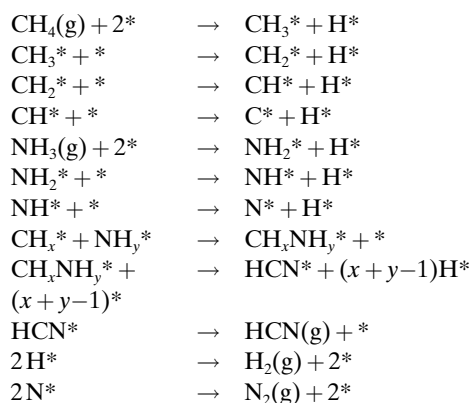
Linear scaling relations have been established between the adsorption energies of hydrogen-containing molecules such as CH_x, NH_x, OH_x, and SH_x and the C, N, O, and S adsorption energies on transition-metal surfaces.^[18] Transition-state energies have also been shown to scale linearly with adsorption energies in a similar fashion.^[9,16,19,20] Recently, a single transition-state scaling relation has been identified for a large number of C–C, C–O, C–N, N–O, N–N, and O–O coupling reactions.^[21] The scaling relations provide a powerful tool for the investigation of reaction mechanisms and the prediction of potential energy surfaces. They limit the number of independent variables to a few, typically adsorption energies of key atoms. Using this information as input to a microkinetic model provides an understanding of trends in catalytic activity. In most cases a volcano-shaped relation

between activity and the key variables, the descriptors, is observed.^[16]

Herein, we provide an example of the approach outlined above and show how one can obtain an understanding of activity/selectivity trends of a reaction with just a few new calculations. We take the synthesis of hydrogen cyanide (HCN), a versatile synthetic building block in organic chemistry, as an example.^[22] For the catalytic production of HCN from methane (CH₄) and ammonia (NH₃) there exist two main processes: the Andrussow and the Degussa (BMA) processes.^[23–25] Both of them use Pt gauze catalysts and are run at temperatures up to 1573 K. The main reason for the high reaction temperature is the endothermicity ($\Delta H^\circ = 251 \text{ kJ mol}^{-1}(\text{HCN})$) of the reaction $\text{CH}_4 + \text{NH}_3 \rightarrow \text{HCN} + 3\text{H}_2$, whereas the competing side reaction $2\text{NH}_3 \rightarrow \text{N}_2 + 3\text{H}_2$ is less endothermic ($\Delta H^\circ = 46 \text{ kJ mol}^{-1}(\text{NH}_3)$). Hence, NH₃ decomposition is thermodynamically favored at low temperatures.

The industrially used Pt/Rh gauze catalysts reach their optimal activity only after an extended activation period under reaction conditions.^[23] SEM^[26] and STM^[27] characterization of the catalyst have shown that during this activation period the catalyst's surface area is increased (roughening) and random faceting occurs. As an approximation of the active site on this surface, we use the stepped (211) facet, which has also been shown to dominate the catalytic activity in other reactions.^[28] Pt is the active catalytic component of the Pt/Rh catalyst, while the role of Rh is to increase the long-term stability.^[23]

We consider the following elementary steps, including nine different routes for C–N coupling between CH_x and NH_y with $x, y = 0, 1, 2$, where X* denotes an adsorbed species and * represents an empty surface site:



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In the present case, we can model the energetics of all CH_x and NH_y species on transition-metal surfaces with the adsorption energies of carbon E_C and nitrogen E_N by using previously discovered scaling relations.^[18] Hydrogen adsorption is modeled with E_C as previously reported.^[29] Thus, the stability of a large fraction of intermediates is given from the literature, and the only missing information is the adsorption energy of CH_xNH_y species. As we will show here, trends in the stability of CH_xNH_y intermediates can be obtained from the energies of the CH_x and NH_y constituents involved in the coupling reaction. Scaling of CH_xNH_y versus E_C and E_N can be found in Figure S1 and Table S1 in the Supporting Information.

The dehydrogenation barriers for NH_3 and CH_4 decomposition follow scaling (or BEP) relations taken from Ref. [30]. Similarly, activation barriers for the C–N and N–N coupling reactions have also been described recently in terms of scaling relations, which link the transition-state energy to the energy of the dissociated products.^[21] The dissociated products can then be linked to E_C and E_N through the scaling relations of adsorption energies of CH_x and NH_x intermediates. Assuming that the C–N coupling reactions are slow when compared to the subsequent hydrogen-exchange reaction of the CH_xNH_y backbone, the thermodynamics and kinetics of the HCN synthesis process are described entirely by E_C and E_N , which leads to a tremendous simplification of the description.

In the following we use this descriptor-based meta-data to study trends in activity and selectivity of the HCN synthesis reaction by coupling it to a microkinetic model and analyze the results on the basis of volcano-shaped relations. For the parameterization of the microkinetic model we corrected the electronic energies obtained from the scaling relations with the zero-point energy E_{ZPE} , heat capacity contributions $\Delta U(T)$, and entropy, because the temperature range of interest for HCN synthesis is quite high. E_{ZPE} , ΔU , and S for each species and all transition states were calculated from the vibrational frequencies on Pt(211) at 1200 K and assumed to be constant. For gas-phase species we used the Shomate equation to estimate the temperature correction and entropy contributions to the Gibbs free energy.^[31] Steady-state numerical solutions of the microkinetic model were obtained under representative reaction conditions for industrial HCN synthesis: $P = 1$ atm, $T = 1373$ K with the initial mole fractions $y_{\text{NH}_3}^0 = 0.107$, $y_{\text{CH}_4}^0 = 0.093$ and assuming a conversion level of 15% with respect to CH_4 and a selectivity of $\text{HCN}/\text{N}_2 = 2$. Further details of the microkinetic model are given in section S2 of the Supporting Information.

By solving this comprehensive microkinetic model, which allows for multiple reaction pathways and the simultaneous production of HCN and N_2 , we obtain a volcano for the overall NH_3 consumption as a function of E_C and E_N as displayed in Figure 1a. The position of several (211) transition-metal surfaces is indicated along with 0.2 eV error bars for E_C and E_N . None of the transition-metal surfaces included here lies on the top of the volcano, but we find Co, Ir, and Ni closest to the top. Rh and Pt are also predicted to be active. Since NH_3 adsorption leads to both HCN and N_2 , it is further important to consider the selectivity shown in Figure 1b. Here

we notice that Pt is unique in terms of its selectivity towards HCN production and is the only metal of those considered that produces only HCN.

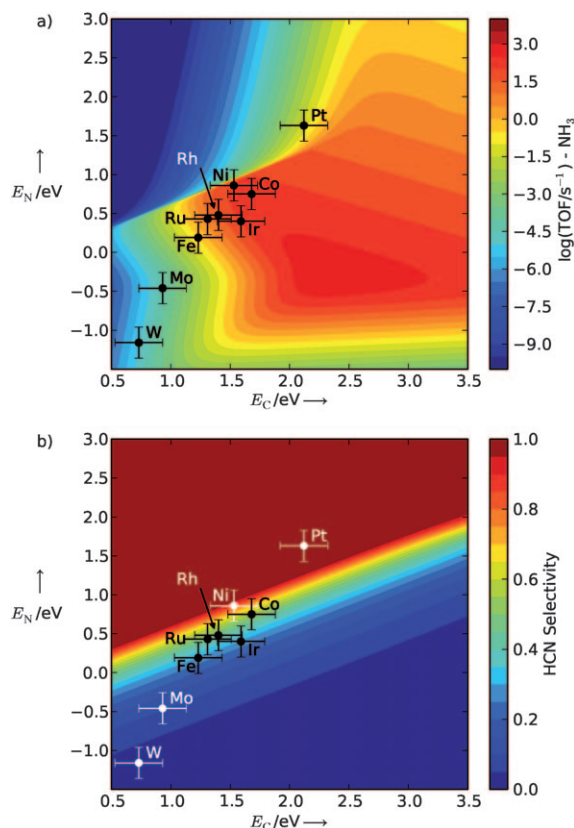


Figure 1. a) Calculated logarithmic turnover frequency (TOF) for NH_3 consumption and b) selectivity towards HCN production as a function of the carbon (E_C) and nitrogen (E_N) binding energies. The error bars indicate an uncertainty of 0.2 eV for E_C and E_N .

To elucidate the competition between HCN and N_2 formation further, we show the HCN and N_2 formation volcanoes separately in Figure 2. Again, we find Co, Ir, and Ni as the most active materials for HCN formation, but at the same time they also produce significant amounts of N_2 (note that the N_2 volcano is one order of magnitude higher in activity). Pt, on the other hand, has a reasonably high activity for HCN synthesis and at the same time does not produce any N_2 . The comparison of the two volcanoes in Figure 2 clearly shows that any material close to the top of the HCN volcano also produces N_2 , which results in a low selectivity. The most interesting region for highly selective and active materials is at weaker E_C and E_N values relative to Pt. In this region the unique selectivity of Pt can be maintained and at the same time higher activities can be achieved.

The mechanism of HCN synthesis has been subjected to numerous studies, but an agreement on the dominant pathway under industrial conditions has not been achieved. Hasenberg and Schmidt studied HCN synthesis from NH_3 and CH_4 on Rh^[32] and Pt,^[33] and also in the presence of O_2 on Pt^[34] at industrially relevant temperatures and low total pressures between 0.01 and 10 Torr. In the absence of O_2 the

reactive Rh and Pt surfaces are covered with almost one monolayer (1 ML) of carbon, which reacts with NH_3 fragments to give HCN. Multiple carbon layers deactivate the catalyst, which provides evidence for a metal-catalyzed reaction where the C–N coupling may occur between differ-

To answer the mechanistic question we inspected the surface coverages and rates of the various C–N coupling reactions. Figure 3 shows that the most abundant surface intermediate for all the investigated metals is carbon. On the very strong binding metals W, Mo, and Fe, the carbon

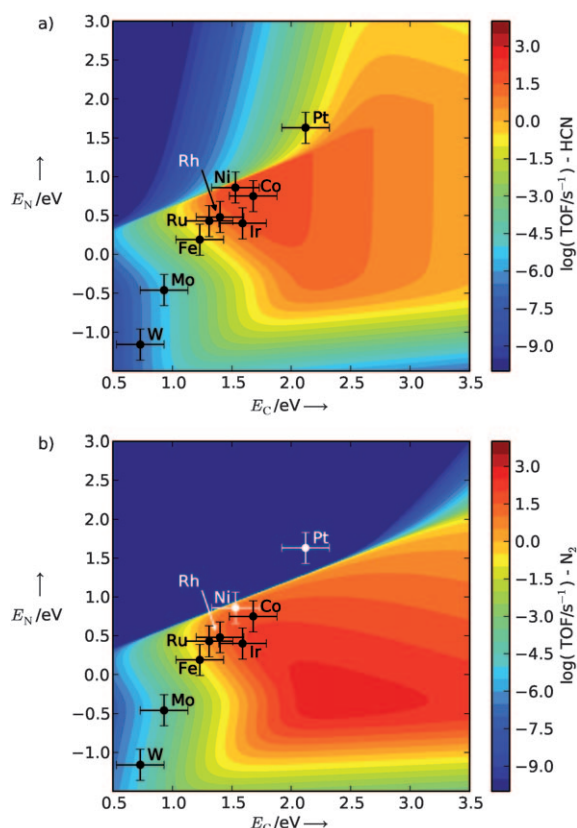


Figure 2. Calculated logarithmic TOF for a) HCN production and b) N_2 production as function of the carbon (E_C) and nitrogen (E_N) binding energies. The error bars indicate an uncertainty of 0.2 eV for E_C and E_N .

ent CH_x and NH_y species under industrial conditions.^[32,33] Evidence for a direct coupling of atomic C and N on the Pt(111) surface from NH_3 and CH_4 at approximately 500 K is provided by ultrahigh-vacuum studies using temperature-programmed desorption (TPD), X-ray photoelectron spectroscopy (XPS), and reflection–absorption IR spectroscopy (RAIRS) on Pt(111).^[35,36] Based on experiments in a temporal analysis of products (TAP) reactor on Pt black at 1173 K, Delagrè and Schuurman^[37] suggested an HCN synthesis reaction mechanism in which NH_3 decomposition is the rate-limiting step and HCN is formed by a C–N or HC–N coupling reaction. However, the TAP experiments do not allow the exclusion of other CH_x – NH_y coupling mechanisms. A very detailed DFT study on Pt(111) suggests that the C–N coupling step in the Degussa process occurs between C or CH and NH_2 .^[38] Lastly, early coupling reactions of CH_x with NH_y on the catalyst surface and subsequent dehydrogenation as a homogeneous gas-phase reaction have also been discussed.^[39,40]

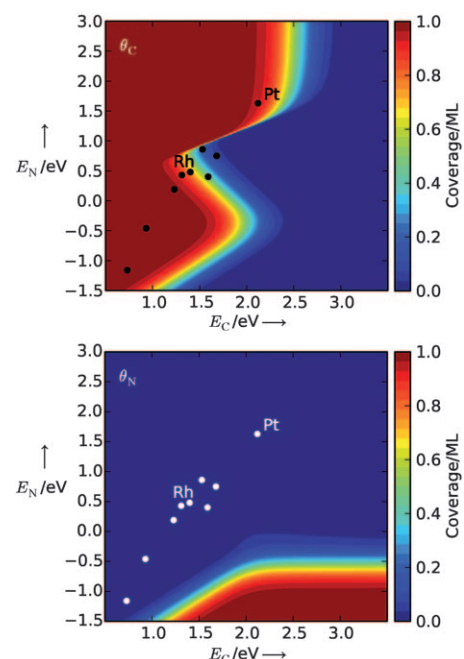


Figure 3. Calculated coverages of C (top) and N (bottom) as a function of the carbon (E_C) and nitrogen (E_N) binding energies. Error bars and metal labels except Pt and Rh have been omitted for clarity.

coverage is 1 ML, whereas for all other metals $\theta_C < 0.95$ ML, which corresponds to the coverage on Pt. This result is in good agreement with the TPD and atomic-emission spectrometry (AES) measurements by Hasenberg and Schmidt.^[32,33] The only other surface species that can be observed in notable quantities is N. High N coverages are predicted for very strong E_N and weak E_C but none of the metals studied herein falls into this region. High surface coverages can alter the stabilities of surface intermediates and transition states and usually lead to a destabilization through repulsive interactions. As has been shown for CO oxidation, one can estimate the effect of adsorbate–adsorbate interactions to a first approximation by moving the metal points to weaker binding energies.^[41] Across different transition-metal surfaces these interactions have a similar effect on the stability of surface intermediates and change the quantitative results, but the observed trends are expected to remain identical.

The individual rates of each possible CH_x and NH_y species are presented as a grid in Figure 4. The panels from left to right in the first row show the coupling rates of C with N, NH , and NH_2 , respectively. Similarly, the first column from top to bottom shows the coupling rates of N with C, CH, and CH_2 . Although it appears as if the direct C–N coupling mechanism is most active in a large region of the investigated parameter space, we note that for Pt the C– NH_2 coupling mechanism is dominant and C–N coupling can be neglected. In contrast, the

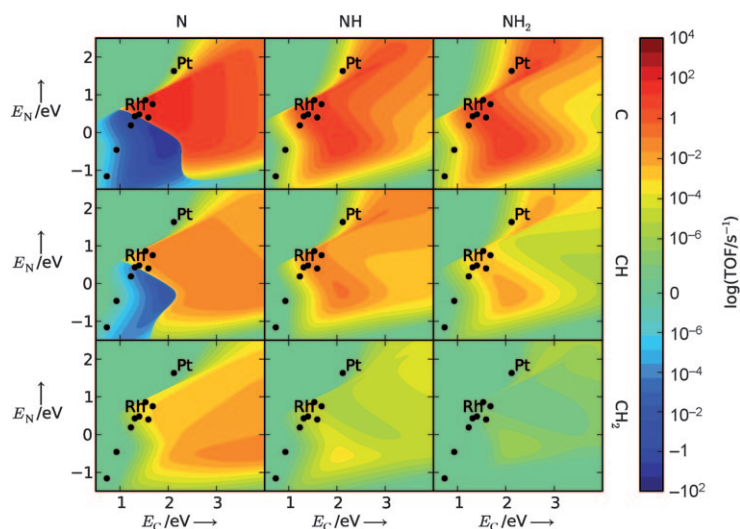


Figure 4. Calculated logarithmic TOFs of C–N coupling steps as a function of the carbon (E_C) and nitrogen (E_N) binding energies. Error bars and metal labels except Pt and Rh have been omitted for clarity. The blue regions indicate negative rates for C–N and HC–N coupling (C–N bond scission).

C–N coupling channel is very active on Rh for which we predict an HCN selectivity of less than 50%. This may explain the high selectivity of Pt towards HCN: an early coupling of NH_y species with surface carbon inhibits the formation of N atoms on the surface, which are necessary for the recombinative desorption as N_2 . Note that the blue regions in the C–N and HC–N coupling graphs indicate negative rates and correspond to C–N bond scission. For materials in this region, NH_3 decomposition to N_2 is facilitated by adsorbed carbon with HCN as a reaction intermediate.

For Pt and Rh, the main components of the industrial catalyst, we also performed a rate-limiting step analysis using the degree of rate control X_{RC} suggested by Campbell.^[42] On Pt the dissociative adsorption of NH_3 is rate-limiting and CH_4 adsorption has a negative impact on the HCN synthesis rate, thereby indicating a carbon poisoning situation. The adverse effect of very high θ_C is also reflected in the negative CH_4 reaction order of -4.3 on Pt. On Rh significant N_2 formation is observed and the mechanism here is more interesting. As on Pt, the adsorption of NH_3 is overall the slowest step in the mechanism. The selectivity is governed by CH_4 adsorption: fast CH_4 adsorption favors the formation of HCN (positive X_{RC}) while N_2 formation is hindered (negative X_{RC}). The X_{RC} of N_2 desorption is negative for both HCN and N_2 formation, which is a counterintuitive result. The explanation for this can be given after a more detailed analysis of the changes in individual rates and coverages. Faster N_2 desorption leads to lower θ_{NH_x} , which causes a slower removal of carbon and increases θ_C . The net effect is a smaller fraction of empty sites and an additional decrease of the NH_3 adsorption rate.

Our results agree well with the recent results by Gómez-Díaz and López^[38] and a number of observations reported for low-pressure HCN synthesis experiments by Hasenberg and Schmidt and the TAP reactor studies by Delagrangé and Schuurman: 1) CH_4 decomposition is fast, 2) the active Pt surface exhibits a high carbon coverage, 3) NH_3 decomposi-

tion is inhibited by surface carbon, and 4) N_2 and HCN formation have NH_3 decomposition as the common rate-determining step. Neither of these experiments, however, can conclusively identify the main C–N coupling step and provide an explanation for the high HCN selectivity of Pt. Delagrangé and Schuurman suggested $\text{C}^* + \text{N}^* \rightarrow \text{CN}^*$ or $\text{CH}^* + \text{N}^* \rightarrow \text{HCN}^*$ as the main routes, but explicitly included the possibility of NH_y^* reactions with C^* . They argue that on Pt these two steps must be faster than N^* recombination to N_2 in order to explain the high selectivity of Pt. Hasenberg and Schmidt simply state that N^* desorbs preferentially as HCN and not as N_2 when CH_4 is present. Our results clearly indicate that on Pt, the step $\text{C}^* + \text{NH}_2^* \rightarrow \text{CNH}_2^*$ leads to the formation of HCN and limits the amount of N^* , which is necessary for N_2 desorption. Based on our results, we propose that the best HCN catalysts in terms of activity and selectivity can be found on the C– NH_2 bond-formation volcano in the little peak area to the right of Pt in Figure 4 (top-right panel).

In summary, we have demonstrated the strength of scaling relations in a descriptor-based approach to determine the reaction mechanism and develop an activity/selectivity map for transition metals based on two parameters for a large-scale heterogeneously catalyzed process. From previously published scaling relations we found E_C and E_N as descriptors that allow the prediction of stabilities and barriers for almost all reaction intermediates and elementary steps in the reaction network. We used these relations in a comprehensive microkinetic model, without making any assumptions regarding the rate-determining step or surface coverages, and constructed volcanoes that not only predict the activity, but also address the issue of selectivity for a wide range of transition-metal surfaces. These volcanoes allow fast screening for new leads for materials with improved catalytic properties. Lastly, we identified the dominant reaction pathway for various possible C–N coupling reactions in this complex reaction network.

We propose the steps outlined above as a general strategy to identify promising transition-metal-based materials/alloys for a wide range of heterogeneously catalyzed reactions. Using the existing universal scaling relations,^[18,21,30] the method allows the rapid discovery of promising candidates without the need for time-consuming/expensive density functional theory (DFT) calculations. Since these universal relations have a relatively large error, we suggest the use of individual scaling lines for the reaction intermediates and transition states in question whenever possible. Since both scaling relations for intermediates and transition states also exist for other classes of materials,^[43] this approach can be applied widely and should not be limited to transition-metal surfaces.

Methods

DFT calculations for the surface-bound CH_xNH_y intermediates were carried out on the (211) surfaces of Ag, Cu, Pt, and Rh using Dacapo. The computational setup for these DFT calculations is identical to

that in previous publications and all details are described in Ref. [21]. The energies reported in this work are given with respect to gas-phase CH_4 , NH_3 , and H_2 .

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- [1] J. K. Nørskov, T. Bligaard, J. Rossmeisl, C. H. Christensen, *Nat. Chem.* **2009**, *1*, 37.
- [2] G. Somorjai, C. Kliewer, *React. Kinet. Catal. Lett.* **2009**, *96*, 191.
- [3] M. P. Andersson, T. Bligaard, A. Kustov, K. E. Larsen, J. Greeley, T. Johannessen, C. H. Christensen, J. K. Nørskov, *J. Catal.* **2006**, *239*, 501.
- [4] J. Sehested, K. Larsen, A. Kustov, A. Frey, T. Johannessen, T. Bligaard, M. Andersson, J. Nørskov, C. Christensen, *Top. Catal.* **2007**, *45*, 9.
- [5] F. Studt, F. Abild-Pedersen, T. Bligaard, R. Z. Sorensen, C. H. Christensen, J. K. Nørskov, *Science* **2008**, *320*, 1320.
- [6] S. Linic, J. Jankowiak, M. A. Barteau, *J. Catal.* **2004**, *224*, 489.
- [7] B. Hinnemann, P. G. Moses, J. Bonde, K. P. Jorgensen, J. H. Nielsen, S. Horch, I. Chorkendorff, J. K. Nørskov, *J. Am. Chem. Soc.* **2005**, *127*, 5308.
- [8] T. F. Jaramillo, K. P. Jorgensen, J. Bonde, J. H. Nielsen, S. Horch, I. Chorkendorff, *Science* **2007**, *317*, 100.
- [9] J. K. Nørskov, T. Bligaard, A. Logadottir, S. Bahn, L. B. Hansen, M. Bollinger, H. Bengard, B. Hammer, Z. Sljivancanin, M. Mavrikakis, Y. Xu, S. Dahl, C. J. H. Jacobsen, *J. Catal.* **2002**, *209*, 275.
- [10] J. K. Nørskov, F. Abild-Pedersen, F. Studt, T. Bligaard, *Proc. Natl. Acad. Sci. USA* **2011**, *108*, 937.
- [11] B. Hammer, J. K. Nørskov, *Surf. Sci.* **1995**, *343*, 211.
- [12] J. Greeley, M. Mavrikakis, *Nat. Mater.* **2004**, *3*, 810.
- [13] V. R. Stamenkovic, B. S. Mun, M. Arenz, K. J. J. Mayrhofer, C. A. Lucas, G. Wang, P. N. Ross, N. M. Markovic, *Nat. Mater.* **2007**, *6*, 241.
- [14] P. Ferrin, A. U. Nilekar, J. Greeley, M. Mavrikakis, J. Rossmeisl, *Surf. Sci.* **2008**, *602*, 3424.
- [15] M. P. Humbert, J. G. Chen, *J. Catal.* **2008**, *257*, 297.
- [16] T. Bligaard, J. K. Nørskov, S. Dahl, J. Matthiesen, C. H. Christensen, J. Sehested, *J. Catal.* **2004**, *224*, 206.
- [17] J. R. Kitchin, J. K. Nørskov, M. A. Barteau, J. G. Chen, *J. Chem. Phys.* **2004**, *120*, 10240.
- [18] F. Abild-Pedersen, J. Greeley, F. Studt, J. Rossmeisl, T. R. Munter, P. G. Moses, E. Skúlason, T. Bligaard, J. K. Nørskov, *Phys. Rev. Lett.* **2007**, *99*, 016105.
- [19] A. Michaelides, Z.-P. Liu, C. J. Zhang, A. Alavi, D. A. King, P. Hu, *J. Am. Chem. Soc.* **2003**, *125*, 3704.
- [20] V. Pallassana, M. Neurock, *J. Catal.* **2000**, *191*, 301.
- [21] S. Wang, B. Temel, G. Jones, L. C. Grabow, F. Studt, T. Bligaard, F. Abild-Pedersen, C. H. Christensen, J. K. Nørskov, *Catal. Lett.* **2011**, *141*, 370.
- [22] K. Weissmehl, J.-H. Arpe, *Industrial Organic Chemistry*, 4th ed., Wiley-VCH, Weinheim, **2003**.
- [23] J. Sauer, M. Bewersdorf, M. Köstner, M. Rinner, D. Wolf in *Handbook of Heterogeneous Catalysis*, 2nd ed. (Eds.: G. Ertl, H. Knözinger, F. Schüth, J. Weitkamp), Wiley-VCH, Weinheim, **2008**, p. 2592.
- [24] L. Andrusow, *Chem. Ing. Tech.* **1955**, *27*, 469.
- [25] E. Koberstein, *Ind. Eng. Chem. Process Des. Dev.* **1973**, *12*, 444.
- [26] L. D. Schmidt, D. Luss, *J. Catal.* **1971**, *22*, 269.
- [27] B. A. Cowans, K. A. Jurman, W. N. Delgass, Y. Z. Li, R. Reifenger, T. A. Koch, *J. Catal.* **1990**, *125*, 501.
- [28] J. Rostrup-Nielsen, J. Nørskov, *Top. Catal.* **2006**, *40*, 45.
- [29] G. Jones, J. G. Jakobsen, S. S. Shim, J. Kleis, M. P. Andersson, J. Rossmeisl, F. Abild-Pedersen, T. Bligaard, S. Helveg, B. Hinnemann, J. R. Rostrup-Nielsen, I. Chorkendorff, J. Sehested, J. K. Nørskov, *J. Catal.* **2008**, *259*, 147.
- [30] S. Wang, V. Petzold, V. Tripkovic, J. Kleis, J. G. Howalt, E. Skúlason, E. Fernández, B. Hvolbæk, G. Jones, A. Toftelund, H. Falsig, M. Björketun, F. Studt, F. Abild-Pedersen, J. Rossmeisl, J. K. Nørskov, T. Bligaard, unpublished results.
- [31] <http://webbook.nist.gov/chemistry/>.
- [32] D. Hasenberg, L. D. Schmidt, *J. Catal.* **1985**, *91*, 116.
- [33] D. Hasenberg, L. D. Schmidt, *J. Catal.* **1986**, *97*, 156.
- [34] D. Hasenberg, L. D. Schmidt, *J. Catal.* **1987**, *104*, 441.
- [35] E. Herceg, M. Trenary, *J. Am. Chem. Soc.* **2003**, *125*, 15758.
- [36] E. Herceg, M. Trenary, *J. Phys. Chem. B* **2005**, *109*, 17560.
- [37] S. Delagrangé, Y. Schuurman, *Catal. Today* **2007**, *121*, 204.
- [38] J. Gómez-Díaz, N. López, *J. Phys. Chem. C* **2011**, *115*, 5667.
- [39] M. Diefenbach, M. Bronstrup, M. Aschi, D. Schroder, H. Schwarz, *J. Am. Chem. Soc.* **1999**, *121*, 10614.
- [40] R. Horn, G. Mestl, M. Thiede, F. C. Jentoft, P. M. Schmidt, M. Bewersdorf, R. Weber, R. Schlogl, *Phys. Chem. Chem. Phys.* **2004**, *6*, 4514.
- [41] L. C. Grabow, B. Hvolbæk, J. K. Nørskov, *Top. Catal.* **2009**, *53*, 298.
- [42] C. T. Campbell, *Top. Catal.* **1994**, *1*, 353.
- [43] E. M. Fernández, P. G. Moses, A. Toftelund, H. A. Hansen, J. I. Martínez, F. Abild-Pedersen, J. Kleis, B. Hinnemann, J. Rossmeisl, T. Bligaard, J. K. Nørskov, *Angew. Chem.* **2008**, *120*, 4761; *Angew. Chem. Int. Ed.* **2008**, *47*, 4683.